**Karan Arora** **R.L. Chemistry Classes M: 99968-68554**

**Max Time : 1 hr** **Class = 11th Chemistry Test**  **Max Marks : 20**

**STRUCTURE OF ATOMS**

**[Upto 3rd Assignment]**

1. Draw a graph between frequency and kinetic energy of ejected electrons. [ 1 ]
2. Draw a graph between frequency and number of ejected electrons. [ 1 ]
3. Neutrons can be found in all atomic nuclei except in one case. What is this atomic nuclei and what does it consist of? [ 1 ]
4. Find out the Atomic number, Mass number , number of Electrons, Proton and Neutrons present in the element with notation 92U238. [ 2 ]
5. Calculate the percentage of higher isotope of neon which has atomic mass 20.2 and the isotope have the mass number 20 and 22. [ 2 ]
6. Find (a) The total number of Electrons (b) The total mass of neutrons in 7mg of C14 . [ 2 ]

(Mass of neutrons is = 1.675 x 10 – 27 kg).

1. Explain electromagnetic wave theory. [ 2 ]
2. In the ultraviolet region of the atomic spectrum of hydrogen, a line is obtained at 1026 . Calculate the energy of the photons of this wavelength. [ 2 ]
3. In the infrared region of atomic spectrum of hydrogen, a line is obtained at 3802 cm – 1. Calculate the energy of this photon. [ 2 ]
4. Write the properties of anode rays. [ 2 ]
5. Explain photoelectric effect. [ 3 ]